

## CHARACTERISTIC STUDIES OF MICRON ZINC PARTICLE HYDROLYSIS IN A FIXED BED REACTOR

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### ABSTRACT

*Zinc fuel is considered as a kind of promising energy sources for marine propeller. As one of the key steps for zinc marine energy power system, zinc hydrolysis process had been studied experimentally in a fixed bed reactor. In this study, we focus on the characteristics of micron zinc particle hydrolysis. The experimental results suggested that the steam inner diffusion is the controlling step of accumulative zinc particles hydrolysis reaction at a relative lower temperature and a relative higher water partial pressure. In other conditions, the chemical reaction kinetics was the controlling step. And two kinds of chemical reaction kinetics appeared in experiments: the surface reaction and the gas-gas reaction. The latter one occurs usually for larger zinc particles and high reaction temperature. Temperature seems to be one of the most important parameters for the dividing of different reaction mechanisms. Several parameters of the hydrolysis process including heating rate, water partial pressure, the particle size and temperature were also studied in this paper. Results show that the initial reaction temperature of zinc hydrolysis in fixed bed is about 410 oC. And the initial reaction temperature increases as the heating rate increases and as the water partial pressure decreases. The total hydrogen yield increases as the heating rate decreases, as the water partial pressure increases, as the zinc particle size decreases, and as the reaction temperature increases. A hydrogen yield of more than 81.5% was obtained in the fixed bed experiments.*

**Keywords:** Hydrogen; Zinc Hydrolysis; Thermal Chemistry; Fixed bed

### INTRODUCTION

Metal fuel is considered as a kind of promising energy sources for marine propeller, especially for military application [1]. Compared to conventional gaseous and liquid fuel, metal fuel usually has very high volumetric energy densities. And the marine energy power system based on the metal fuel can use ocean water as part of propellant oxidizing agent directly, which can save a great part of space and weight for fuel carrier underwater.

Zinc is considered as a kind of promising candidate of metal fuel for marine energy power system. It has been widely studied for power generation via the two-step thermochemical cycle based on Zn/ZnO redox pair for its brilliant hydrolysis ability [2,3,4]. The Schematic diagram of the marine energy power system based on zinc fuel is shown in Fig.1. Firstly ocean water

is introduced into the first combustion room, where the zinc powder fuel hydrolysis with the water and the hydrogen gas is produced at a relative low temperature. This reaction is a heat releasing reaction, and the reactant can be heated to reaction temperature automatically [5]. Secondary, the produced hydrogen is introduced into the second combustion room, where the hydrogen combusts with oxidant and release great amount of heat at a very high temperature. The residual ocean water in combustion room is heated to be vapor and then be sprayed out with the final combustion products.

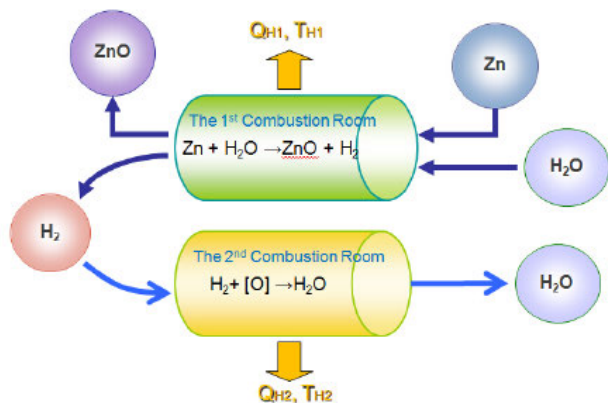


Fig. 1. Schematic diagram of the marine energy power system based on zinc fuel.

As one of the key steps for marine energy power system based on zinc fuel, several studies were carried out on the zinc hydrolysis reaction. In a series of thermogravimetric analyses of commercial zinc powder and solar zinc powder in a temperature range of 350-500 °C, Weidenkaff et al. [3] found that the hydrolysis reaction proceeded faster for molten zinc and for zinc containing impurities, but a layer of ZnO prevented the reaction from reaching completion. The hydrolysis of submicron Zn particles in a temperature range of 330-360 °C was also studied by thermogravimetric analyses, and a fast surface reaction, corresponding to a mass increase of 2%, followed by a slow diffusion-limited reaction was observed [5]. The oxidation of liquid zinc with water vapor was studied by bubbling water vapor through bulk of liquid zinc at 450-500 °C [6], and the results showed that the specific reaction rate increases as the water partial pressure increases, the main determining step of the hydrolysis reaction is the diffusion of reactants through the product zinc oxide layer. The oxidation of zinc vapor of about 750°C and 800 °C with water vapor was also studied using a tubular aerosol flow reactor which features three temperature-controlled zones [7], up to 83% of zinc conversion could be obtained while the temperature of reaction zone is just below the Zn(g) saturation temperature. And several reaction parameters of the H<sub>2</sub> production by steam-quenching of Zn vapor were also studied in a hot-wall aerosol flow reactor, the results shows that high zinc conversions could get at a low quenching rate at the expense of low particle yield [8].

Detailed studies on the hydrolysis kinetics of zinc powder was also conducted by Vishnevetsky et al. [9], they found that the hydrolysis of zinc proceeded in two stage, and increasing the beginning temperature of the reaction is advantageous to the hydrolysis process, the reactivity of Solar Zinc is much higher than commercial zinc, which is agreed with paper [3]. Thermogravimetric analysis of the hydrolysis of zinc particles was also take out in a temperature range of 200-1000°C [10], two kinds of reaction mechanism of the zinc hydrolysis were revealed by Lv et al, in which the rate of hydrolysis reaction was limited by the evaporation of zinc and the diffusion of

zinc through ZnO layer respectively.

The hydrolysis of zinc particle with nano scale was studied by Ma et al [11] and Bhaskar et al [12]. The complete conversion of Zn nano crystals with 70nm average size was achieved at 175°C with a 19% water vapour mole fraction [11]. A hydrogen yield of 80% was achieved at 600 °C for the hydrolysis of a unique Zn nano particle (~25nm) which was dispersed on carbon micro-nano fibers [12].

In this paper, detailed thermodynamic studies were carried out for the hydrolysis process of micron zinc particles in a fixed bed reactor. We tried to Fig. out the characteristics of the zinc hydrolysis process, especially in high water partial pressure. The influences of several important process parameters such as heating rate, reaction temperature, water partial pressure and particle size were also studied here. We focused not only on the final hydrogen yield, but also on the influences on the characteristics of reaction process.

## EXPERIMENTAL

The fixed bed experimental system is depicted schematically in Fig.2. The steam supplied by a steam generator (Shanghai Fengxian, model KQ-BII) was preheated to temperature T<sub>1</sub> and then mixed with the inert carrier gas. The mixed gases were introduced into the reaction zone of a quartz tube reactor (with a inner diameter of 25mm). The quartz tube reactor was placed in the furnace of a high temperature tube furnace (Shanghai Sanya, model SK2-2.5-13TS). A ceramic boat with sample zinc particles inside was placed in the reaction zone of the quartz tube reactor. The exhaust gases from the tube reactor were cooled by a water cooler and dried by a silica gel dryer, and then went through an in situ hydrogen concentration analyzer and a rapid flowmeter FC2 (Swiss Sensrion), and finally be vented into atmosphere.

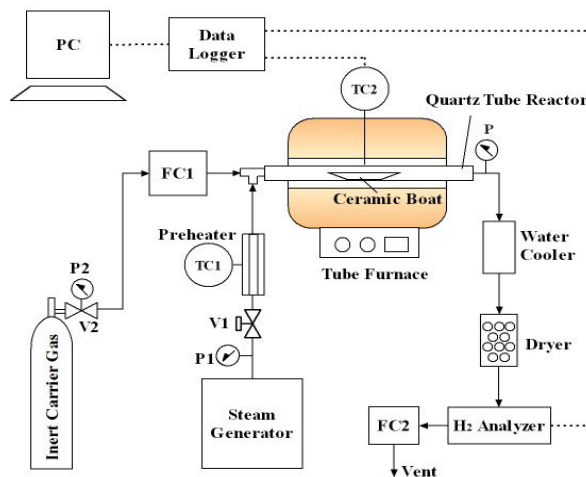


Fig. 2 Schematic diagram of the fixed bed experimental system.

The inert carrier gas used argon gas with a purity of 99.99%. It's flow rate was controlled by a mass flow rate controller FC1 (Qixing Huachuang, model D07). The flow rate of steam was controlled by a needle valve V1. The reaction zone of the quartz tube reactor was heated by the tube furnace, which has 3 Levels

of heating power to choose. The heating curve of each power Level is shown in Fig. 3. As Fig. 3 shows, the heating curve of temperature versus time of the tube furnace follows logarithm curve approximately. And when the heating temperature rises from 200 °C to 900°C, the average heating rates at Level 1 to Level 3 heating power were calculated as 1.17, 2.50 and 4.61 °C/min respectively.

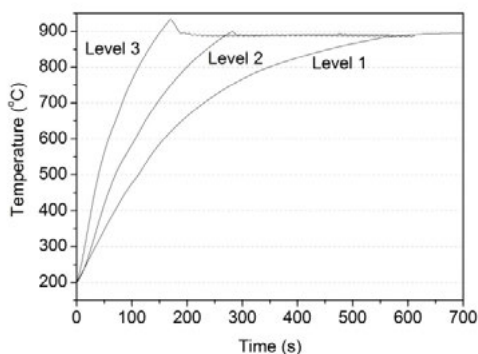


Fig. 3 The heating temperature of tube furnace versus time at 3 Levels of heating power (the initial temperatures were set at 200 °C).

The temperature of reaction zone  $T_2$  was measured by a K type thermocouple TC2, as shown in Fig.2. The in situ hydrogen concentration analyzer (Hitech, model K522) was used to measure the volume concentration of hydrogen in argon-hydrogen mixture. It has a measuring scale of 0-30% and an accuracy of 1% F.S. The pressure of the reaction zone  $p$  was tested in situ by a pressure sensor (Shanghai Tianmu, model NS-F) at the outlet of tube reactor. The data of reaction temperature  $T_2$  and hydrogen concentration were collected continuously and simultaneously by a data logger (Agilent, Model 34970A).

There are two kinds of sample zinc. One is the high purity zinc powder (by Sinopharm Chemical Reagent Co., Ltd) with purity of 99.99%, and its mean particle size is 13  $\mu\text{m}$ . The other kind is the common commercial Zn powder (by Changsha Wellink Zinc Material Co., Ltd) with purity of 99%, and four groups of common Zn powder with different mean particle size (11  $\mu\text{m}$ , 45  $\mu\text{m}$ , 108  $\mu\text{m}$  and 178  $\mu\text{m}$ ) were used in the experiments to discuss the influence of particle size. In experiments, the zinc powder samples were hold by a ceramic boat with a size of 77 mm×12 mm×9 mm. The sample weight  $m_{\text{Zn},i}$  in each experimental case was weigh about 0.50±0.05g. The sample zinc powders were dispersed equally at the bottom of ceramic boat.

In each experimental case, the initial gauge pressure of inert carrier gas was set at 0.05 Mpa; And the flow rate of inert carrier gas was set at 200 ml/min; The initial gauge pressure of steam from steam generator was set at 0.1 Mpa; The temperature  $T_1$  of steam preheater was set at 180-200 °C.

There were two kinds of experimental cases: isothermal cases and non-isothermal cases. The isothermal cases were mainly used to study the influences of reaction temperature. Experimental methods are described as follows:

#### (a) isothermal cases

Firstly the sample was placed at the reaction zone of

the quartz tube reactor. And then argon gas with fixed pressure and flow rate was introduced into tube reactor for 10min to purge out the residual air in tube reactor. Then the heating temperature of tube furnace was heated to the required isothermal temperature. Once the required heating temperature was reached, the valve V1 was open to introduce the steam with fixed temperature, pressure and flow rate into tube reactor. After the completion of reaction (when the in situ scanned hydrogen concentration was observed as zero again), the flow rate of steam was measured.

#### (b) non-isothermal cases

Firstly the sample was placed at the reaction zone of the quartz tube reactor. And then argon gas with fixed pressure and flow rate was introduced into tube reactor for 10min to purge out the residual air in tube reactor. Then the heating temperature of tube furnace was heated from room temperature to 200 °C, and the heating power of tube furnace was set at the required Level. And then the valve V1 was open to introduce the steam with fixed temperature, pressure and flow rate into tube reactor. And the heating temperature was raised from 200 °C to 900 °C under fixed power Level and then be kept at 900 °C. After the completion of reaction, the flow rate of steam was measured.

From the pressure  $p$  measured at the outlet of tube reactor in experiments (Fig.2), it is showed that the reaction pressure in our experimental setup was always kept at atmospheric pressure. The system pressure changed every little during the whole experiments. So it is reasonable to consider the reaction in fixed bed system as an atmospheric reaction.

A dryer was used to measure the mass flow rate of steam  $F_{\text{H}_2\text{O}}$  (mg/min) after each cases. It was weighed before and after each measurement in a fixed drying time, the flow rate of steam was calculated as the weight change of dryer divided by the drying time. As the flow rate of carrier gas (Ar) was constantly set to 200 ml/min, the water partial pressure (bar) in each case can be calculated as follows:

$$p_{\text{H}_2\text{O}} = p_{\text{system}} \times F_{\text{H}_2\text{O}} \times 22.4 / (F_{\text{H}_2\text{O}} \times 22.4 + 18 \times 200)$$

The elements of solid products were analyzed by energy dispersive x-ray spectroscopy (EDX, FEI Model SIRION-100). The morphologies of the solid products were examined by scanning electron microscopy (SEM, FEI Model SIRION-100).

## RESULTS AND DISCUSSION

### TYPICAL SOLID PRODUCTS ANALYSIS

In fixed bed experiments, the solid products of zinc hydrolysis appeared as white or grey solid blocks, as Fig.4 shows. The colour of solid products depends on the conversion of zinc in each case. The SEM pictures of typical solid product were shown in Fig.5. It is shown that the solid product of zinc particle hydrolysis in fixed bed reactor has a porous structure. The EDX analysis results showed that the main element of solid products is zinc oxide. As Fig.5(b) &(d) shows, the porous solid products are constructed by a serials of zinc oxide

spherical shells. The average thickness of spherical shells was measured about 1-2 $\mu\text{m}$ . The average diameter of spherical shells was measured about 10 $\mu\text{m}$ , which is similar to the average particle size of reactant zinc particles. This kind of product structure for hydrolysis of zinc particles was also discovered in thermogravimetric analysis [12]. It is suggested that the hydrolysis of micron pure zinc particles in our experiments is a surface reaction. For surface reaction between gas-solid reactants, the reaction rate is usually controlled by the chemical reaction kinetics or the gas reactants diffusion. There are two kinds of chemical reaction mechanism of zinc particle hydrolysis, which says that the rate of hydrolysis reaction is limited by the evaporation of zinc or the diffusion of zinc through ZnO layer respectively [12]. The former one usually happens for larger particles with a diameters above 100 $\mu\text{m}$  [12]. In our experiments, obviously the latter mechanism is the right chemical reaction mechanism for the hydrolysis of pure zinc particle samples.

In TG experiments just as in literature [12], the sample weight is quite small and the layer thickness of reactant zinc particles accumulation in sample holder is very thin, so the influence of reactant gas diffusion could be ignored. But in fixed bed experiments here, the sample weight was respectively larger, and the average layer thickness of reactant zinc particles accumulation in ceramic boat was measured about 1 mm, which was thick enough to consider the influence of steam diffusion in reaction. In addition, the porous structure of solid products, as Fig.5 (d) shows, confirms the necessity of steam diffusion consideration.

#### TYPICAL HYDROLYSIS PROCESS ANALYSIS IN FIXED BED REACTOR

The typical hydrogen concentration curves in non-isothermal cases were shown in Fig.6. From the Fig. we can see obviously that there are two reaction stages in non-isothermal cases. Each reaction stage corresponds to a peak of hydrogen concentration curve in different temperature ranges. Obviously the control mechanisms of the two stages are different. As analyzed before (see 3.1), the controlling step of zinc particle hydrolysis in fixed bed may be the chemical reaction kinetics or the diffusion of reactant steam. In following discussion we tried to Fig. out the right control mechanism of each reaction stage.

We defined five characteristic points to analyze the hydrolysis process, as shown in Fig. 6(a) &(b). Each point represents a special meaning, just as follows:

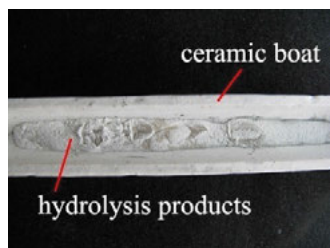


Fig. 4 Photo of typical solid products of zinc hydrolysis in ceramic boat.

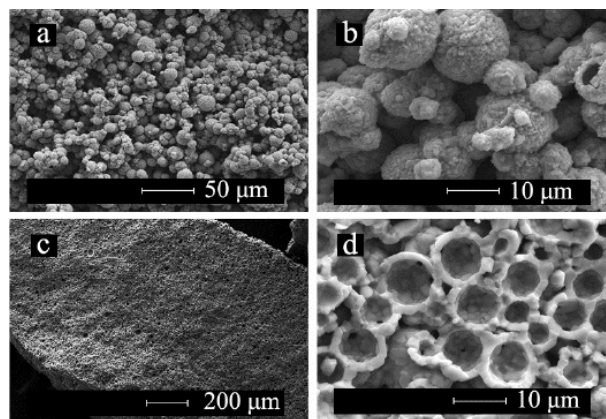


Fig. 5 SEM micrographs of solid products for experiments at a typical parametric setting ( $p=0.1\text{ MPa}$ ,  $p_{\text{H}_2\text{O}} = 0.76\text{ bar}$ , Level 1 heating power, pure zinc powders,  $m_{\text{Zn}}=0.45\text{g}$ ):(a)&(b)the micrographs of the top surface of products;(c)&(d)the micrographs of the cross section of the products.

Point A: The initial reaction point, when hydrogen concentration was observed at the first time in each case. The time of Point A represents the initial reaction time  $t_i$ . And the temperature of Point A represents the initial reaction temperature  $T_i$ .

Point B: The peak hydrogen concentration point in the first stage. The hydrogen concentration of Point B  $n_{\text{max},1}$  is used to represent the maximum reaction rate in the first reaction stage.

Point C: The valley bottom point between two neighbouring peaks. The temperature of Point C  $T_D$  is used to represent the dividing temperature of the two reaction stages.

Point D: The peak hydrogen concentration point in the second stage. The hydrogen concentration of Point D  $n_{\text{max},2}$  is used to represent the maximum reaction rate in the second reaction stage.

Point E: The finish point of reaction, while the hydrogen concentration dropped to zero again. The time of Point E represents the final reaction time  $t_f$ . And the temperature of Point E represents the final reaction temperature  $T_f$ . Comparing the curves in Fig. 6(a) and (b), we can see that the heating temperature went into isothermal stage in the typical case before the end of reaction. So there's no discussion meaning for  $T_f$ .

In each case, the total hydrogen yield  $V_{\text{H}_2}$  can be calculated as follows:

$$V_{\text{H}_2} [L] = \int_{t_i}^{t_f} V(t) dt$$

Where  $V(t)$  is the function of produced hydrogen flow rate versus time. It is given by:

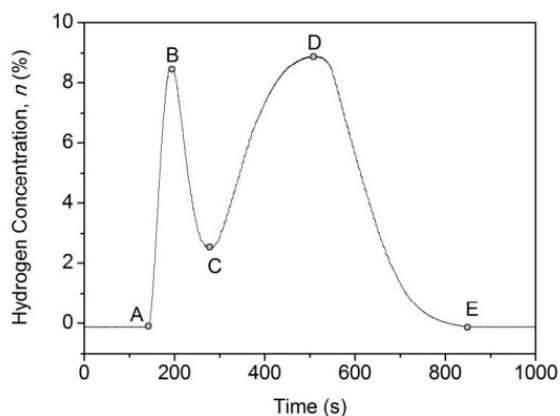
$$V(t) [L/s] = \frac{(V_{\text{Ar}} / 60) \times n}{100 - n}$$

Where  $V_{\text{Ar}}$  is the flow rate of argon carrier gas, L/min;  $n$  is the data of hydrogen concentration measured in situ by the hydrogen analyzer, %.

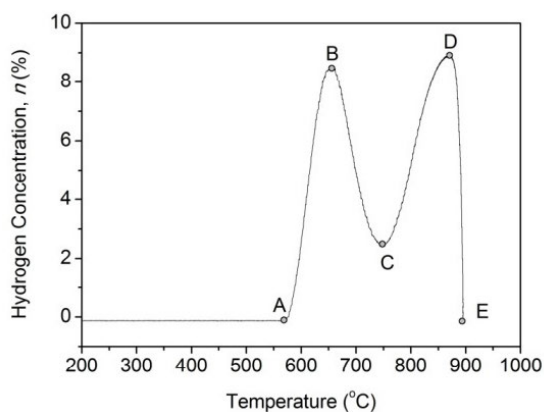
So the total hydrogen yield percentage in each case can be calculated as:

$$H_2 \text{ Yield } [\%] = (V_{H_2} / 22.4) / (m_{Zn,i} / 65)$$

Where  $m_{Zn,i}$  is the initial sample weight of zinc particles, g.



(a) Time curve



(b) Temperature curve

Fig. 6 Typical hydrogen concentration curve in non-isothermal cases ( $p=0.1$  MPa,  $p_{H_2O} = 0.76$  bar, Level 1 heating power, pure zinc powders,  $m_{Zn,i}=0.45$ g).

### THE INFLUENCE OF HEATING RATE

Keeping the water partial pressure as constant approximately, the influences of heating rate were studied. As described before, the average heating rates at Level 1, Level 2 and Level 3 heating power were 1.17, 2.50 and 4.61 °C/min respectively, when the heating temperature rises from 200 °C to 900°C. So the influences of Levels of heating power represent the influences of heating rates in other hand.

Experimental results were shown in Fig.7 and Tab. 1. The two-peak structure of hydrogen concentration curve is more prominent under Level 1 heating power. The initial reaction temperature rises as heating rate rises, as Tab. 1 shows. The  $T_i$  in Level 1 heating power case was about 572°C, while the  $T_i$  in Level 3 heating power case was about 770 °C. The TG studies show that the initial reaction of micron zinc particle hydrolysis is about 600°C [12]. So it is suggested that the non-isothermal experiments carried out under low heating rate with Level 1 heating power can reflect more real reaction characteristics. So the heating powers of tube furnace in the following cases

were all set at Level 1.

The results also show that the hydrogen yield decreases slightly from about 77% to 72% as the heating power increases from Level 1 to Level 3.

Tab.1 Influences of heating power ( $p=0.1$  MPa,  $p_{H_2O} = 0.75-0.76$  bar, pure zinc powders,  $m_{Zn,i}=0.45-0.46$ g).

Heating Power	H <sub>2</sub> Yield / %	n <sub>max,1</sub> / %	n <sub>max,2</sub> / %	T <sub>i</sub> / °C	T <sub>D</sub> / °C
Level 1	76.95	8.53	8.91	572	750
Level 2	75.59	--	17.00	702	--
Level 3	72.27	7.78	16.72	770	--

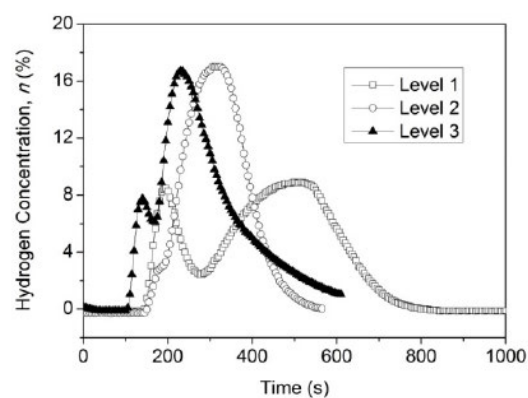


Fig. 7 Hydrogen concentration vs. time at 3 Levels of heating power ( $p=0.1$  MPa,  $p_{H_2O} = 0.75-0.76$  bar, pure zinc powders,  $m_{Zn,i}=0.45-0.46$ g).

### THE INFLUENCE OF WATER PARTIAL PRESSURE

Experiments were carried out to study the influences of water partial pressure in a scale of 0.59 bar to 0.90 bar, as Fig. 8 shows. A water partial pressure as high as 0.90 bar under atmospheric system pressure could be got in our fixed bed experiments. The characteristic parameters of hydrolysis process under different water partial pressures are listed in Tab. 2.

From the experimental results we can see that as the water partial pressure decreases, the hydrogen yield decreases, the initial reaction temperature increases, and the dividing temperature increases either, as Tab. 2 shows. The height of hydrogen concentration peak in the first stage decreases quickly while the water partial pressure decreases. The height of hydrogen concentration peak in the second stage increases as the water partial pressure decreases, just appears opposite tendency to the first stage, as Fig.8 shows. And when the water partial pressure was as low as 0.59 bar, only one peak appears. Considering the temperature range of that peak and the change rule of the peak height versus the water partial pressure, it is reasonable to deduce that the remaining one peak represents the second stage of reaction. In other words, the hydrogen

concentration peak in the first stage disappeared under low water partial pressure.

As mentioned before, there are two kinds of controlling steps of zinc particle hydrolysis in the two reaction stages, the chemical reaction kinetics or the diffusion of reactant steam. According to the general surface reaction theory, the diffusion rate of steam depends very much on the water partial pressure. And as the water partial pressure decreases, the diffusion rate of steam will decrease too. Meanwhile for the chemical reaction kinetics of zinc particles hydrolysis in these cases, the diffusion rate of zinc through ZnO layer depends more on temperature, and little on water partial pressure. Considering that in the first reaction stage with lower temperature range, the hydrogen peak height decreases as the water partial pressure decreases, which has the similar change rule with the diffusion rate of steam versus water partial pressure. Obviously the steam diffusion seems to be the right controlling step of the first reaction stage in lower temperature range. Be more precisely speaking, the steam's diffusing inside the accumulative zinc particles (which can be seen as porous block) is the controlling step in the first reaction stage. It could be called as "inner diffusion" in surface reaction. Meanwhile the chemical reaction kinetics, that the diffusion of zinc through ZnO layer, is suggested to be the controlling step of the second reaction stage in higher temperature range.

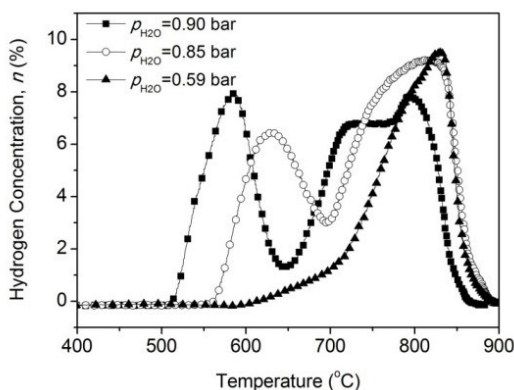


Fig. 8 Hydrogen concentration vs. temperature under different water partial pressures ( $p=0.1$  MPa,  $p_{H_2O} = 0.59-0.90$  bar, Level 1 heating power, pure zinc powders,  $m_{Zn,i}=0.52-0.53$ g).

Tab. 2 Influences of water partial pressure ( $p=0.1$  MPa,  $p_{H_2O} = 0.59-0.90$  bar, Level 1 heating power, pure zinc powders,  $m_{Zn,i}=0.52-0.53$ g).

$p_{H_2O}$ /bar	$H_2$ Yield / %	$n_{max,1}$ / %	$n_{max,2}$ / %	$T_i$ / °C	$T_D$ / °C
0.90	75.98	7.93	7.85	512	653
0.85	70.09	6.43	9.22	558	698
0.59	55.06	0.00	9.53	596	--

## THE INFLUENCE OF ZINC PARTICLE SIZES

Experiments were carried out to study the influences of zinc particle sizes, as shown in Fig. 9 and Fig. 10. The characteristic parameters of hydrolysis process for different size of zinc

particles are listed in Tab. 3.

The experimental results show that the peak hydrogen concentration in the first and the second stages both decrease as the mean particles size increases. And the hydrogen yield dropped from 77.45% to 11.27% when the mean particle size increased from 11 $\mu$ m to 178 $\mu$ m, as Fig.10 shows.

Furthermore, there were three hydrogen concentration peaks appeared in the hydrolysis process of sample Zn-4, which has a mean diameter of 178 $\mu$ m, as Fig.9 shows. The initial temperature of the third reaction stage is about 752 °C. From the past research results of literature [12], the controlling step of the third reaction stage is speculated to be the evaporation of zinc nucleus. That means the chemical reaction mechanism in the third stage for sample Zn-4's hydrolysis is mainly gas-gas reaction, but not surface reaction. And the typical microstructure of solid products in gas-gas zinc hydrolysis reaction was discovered in the SEM pictures of sample Zn-4's hydrolysis products, as Fig.11 shows. There were many small zinc oxide crystal needles covered outside of ZnO spherical shell, as Fig.11(b) shows.

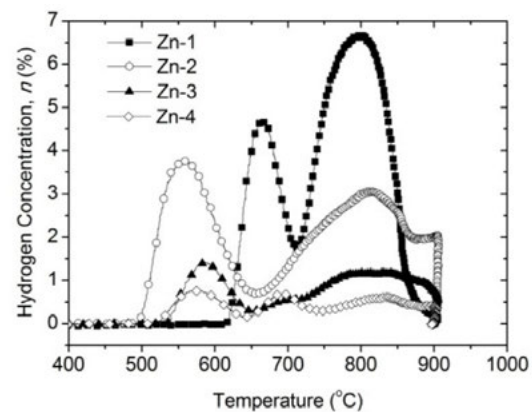


Fig. 9 Hydrogen concentration vs. temperature for different zinc particle sizes ( $p=0.1$  MPa,  $p_{H_2O} = 0.75-0.77$  bar, Level 1 heating power, common commercial zinc powders, the mean particle size for Zn-1, Zn-2, Zn-3 and Zn-4 is 11 $\mu$ m, 45 $\mu$ m, 108 $\mu$ m and 178 $\mu$ m respectively,  $m_{Zn,i}=0.51-0.52$ g).

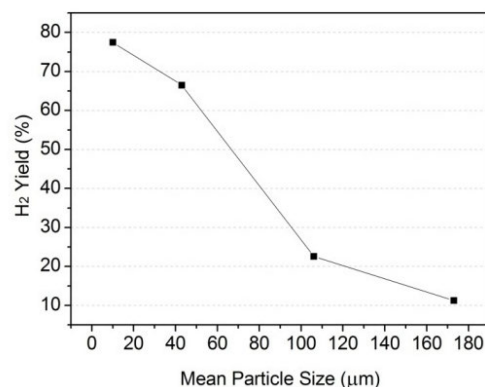


Fig. 10 The influences of particle size on  $H_2$  yield ( $p=0.1$  MPa,  $p_{H_2O} = 0.75-0.77$  bar, Level 1 heating power, common commercial zinc powders,  $m_{Zn,i}=0.51-0.52$ g).

Tab. 3 Influences of zinc particle size ( $p=0.1$  MPa,  $p_{H_2O} = 0.75-0.77$  bar, Level 1 heating power, common commercial zinc powders,  $m_{Zn,i}=0.51-0.52$ g).

Mean particle size / $\mu\text{m}$	H <sub>2</sub> Yield/ %	$n_{\text{max},1}$ / %	$n_{\text{max},2}$ / %	$n_{\text{max},3}$ / %	$T_i$ / °C	$T_{D1}$ / °C
11	77.45	4.69	6.68	--	647	742
45	66.47	3.77	3.06	--	496	661
108	22.54	1.63	1.39	--	534	655
178	11.27	0.77	0.69	0.60	522	645

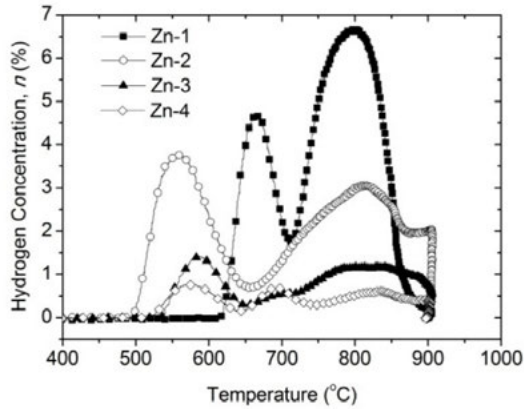


Fig. 11 SEM micrographs of solid products for the hydrolysis of zinc particles with a mean particle size of 178  $\mu\text{m}$  ( $p=0.1$  MPa,  $p_{H_2O} = 0.76$  bar, Level 1 heating power, common commercial zinc powders,  $m_{Zn,i}=0.51$ g).

### THE INFLUENCE OF REACTION TEMPERATURE IN ISOTHERMAL CONDITIONS

A series of isothermal experimental cases were carried out to study the influence of reaction temperature, as Fig.12 shows. In this part of discussion, we focused on the initial reaction temperature of zinc particles. So the temperature scale of isothermal cases was mainly from 400 °C to 600 °C. In addition, the zinc hydrolysis isothermal case with a temperature of 900 °C was carried out as well to study the reaction process near the boiling point of zinc.

The experimental results are shown in Fig.12-Fig.14. It should be pointed out firstly that because of the restriction of heating rate, when the fixed reaction temperature was higher than the melting point of zinc, the zinc sample would be exposed in flowing inert carrier gas for a period of time. Because of the melting and evaporating of zinc particles, there will be a little loss of the zinc sample before the reaction start. Obviously, higher the temperature fixed, larger the zinc sample loss. This is another reason why the most of our isothermal cases carried out in low temperature range.

From Fig.12 we can see that there was no hydrogen produced at a temperature of 400 °C, while the hydrogen peak appeared in the case of 450 °C. That means the initial reaction temperature of zinc particle reaction in fixed bed reactor is between 400-450 °C. A series of isothermal experiments with very small temperature step were carried out. And the results show that the initial reaction temperature of micron zinc particle hydrolysis in fixed bed is about 410 °C.

Be different from the multi-peak structure in non-isothermal cases as mentioned previously, the hydrogen

concentration curve in isothermal cases usually had only one peak, as Fig.12 shows. That suggested that the temperature is one of the most important characters for the reaction mechanism.

When the temperature is lower than 600 °C, the peak hydrogen concentration and hydrogen yield both increase almost linearly as the isothermal temperature increases, as shown in Fig.13 and Fig.14. Obviously at a temperature near the boiling point of zinc such as 900 °C, the hydrolysis of zinc proceeded more well, as the peak hydrogen concentration was as large as 22.9%, and the hydrogen yield was as large as 81.5%. Considering the sample loss before the reaction starting in this case, the real hydrogen yield for 900 °C case should be larger than 81.5%. That may because the gasification of zinc at very high temperature becomes more violently, so the hydrolysis process based on gas-gas reaction proceeds more quickly.

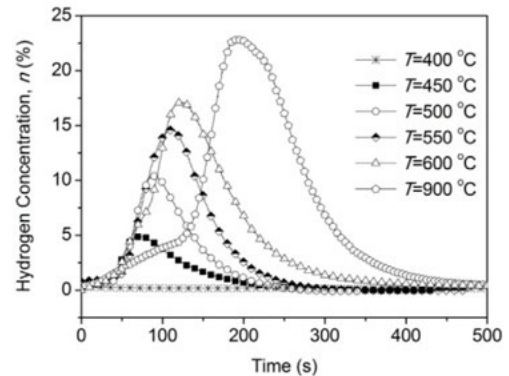


Fig. 12 Hydrogen concentration vs. time under different temperatures in isothermal cases ( $p=0.1$  MPa,  $p_{H_2O} = 0.75-0.77$  bar, Level 1 heating power, pure zinc powders,  $m_{Zn,i}=0.49-0.51$ g).

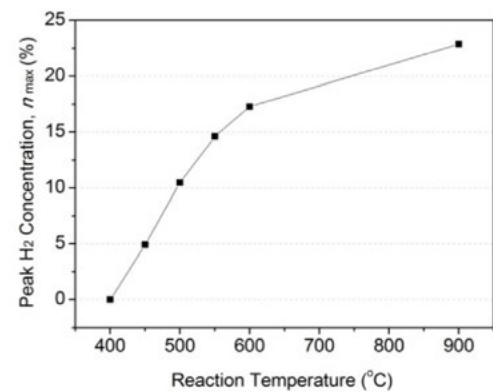


Fig. 13 The influences of temperature on peak hydrogen concentration ( $p=0.1$  MPa,  $p_{H_2O} = 0.75-0.77$  bar, Level 1 heating power, pure zinc powders,  $m_{Zn,i}=0.49-0.51$ g).

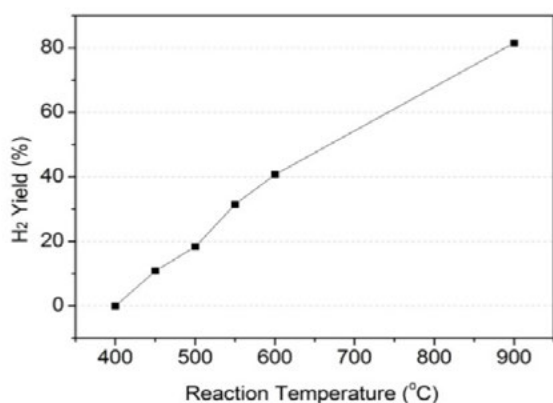


Fig. 14 The influences of temperature on H<sub>2</sub> yield ( $p=0.1$  MPa,  $p_{\text{H}_2\text{O}} = 0.75$ - $0.77$  bar, Level 1 heating power, pure zinc powders,  $m_{\text{Zn}}=0.49$ - $0.51$ g).

## CONCLUSION

A series of experiments in a fixed bed reactor were carried out to study the characteristics of micron zinc particle hydrolysis process. The experimental results suggested that the steam inner diffusion is the controlling step of accumulative zinc particles hydrolysis reaction at a relative lower temperature and a relative higher water partial pressure. In other conditions, the chemical reaction kinetics was the controlling step. And two kinds of chemical reaction kinetics appeared in experiments: the surface reaction and the gas-gas reaction. The latter one occurs usually for larger zinc particles and high reaction temperature. Temperature seems to be one of the most important parameters for the dividing of different reaction mechanisms.

Several parameters of the hydrolysis process including heating rate, water partial pressure, the particle size and temperature were also studied in this paper. Results show that the initial reaction temperature of zinc hydrolysis in fixed bed is about 410 °C. And the initial reaction temperature increases as the heating rate increases and as the water partial pressure decreases. The total hydrogen yield increases as the heating rate decreases, as the water partial pressure increases, as the zinc particle size decreases, and as the reaction temperature increases. A hydrogen yield of more than 81.5% was obtained in the fixed bed experiments.

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